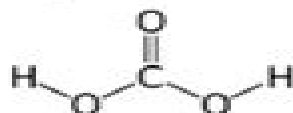


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PERIOD 1:

DEFINITION OF ORGANIC CHEMISTRY

Organic chemistry is the study of the structure, properties, composition, reactions, and preparation of carbon-containing compounds. Most organic compounds contain carbon and hydrogen, but they may also include any number of other elements (e.g., nitrogen, oxygen, halogens, phosphorus, silicon, sulfur). Carbon dioxide (CO_2), water (H_2O), and calcium carbonate (CaCO_3) are not organic compounds because they do not contain carbon-hydrogen (C-H) bonds.

 H_2CO_3 **UNIQUE NATURE OF CARBON**

Carbon atom belongs to group IV of the periodic table of element.

1. The excited state electronic configuration of carbon is $1s^2 2s^1 2p_x^1 2p_y^1 2p_z^1$. A carbon has four electrons in its outer shell and uses by sharing electrons to form covalent bonds (tetravalency).
2. It has no d-orbital available for bonding, therefore it can only form four bonds, and these gives rise to a variety of neutral compounds having single, double and triple bonds.
3. There are three (3) types of orbital hybridization available to carbon namely Sp^3 , Sp^2 and Sp . **Hybridization** is a process in which two or more atomic orbitals of comparable energy in the valence-shell of an atom (central atom of a molecule or ion) mix together and give rise to the formation of new degenerate (having the same energy level) orbitals known as hybrid orbitals. The size of the hybrid orbital decreases as the percentage of s-character increases. As a result, the size of the hybrid orbital is: $\text{sp}^3 > \text{sp}^2 > \text{sp}$.
4. Catenation: Carbon atoms have a unique property called catenation, which allows them to bond with one another in a long and continuous chain, leading to the creation of carbon-based backbones in organic molecules.
5. The bond between carbon and hydrogen is almost non-polar. Thus, hydrogen atoms attached to carbon do not weaken the C-C bonds. The electronegativity value of carbon and hydrogen are 2.5 and 2.1 respectively.

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