



Lab 3: Stoichiometry I

Stoichiometry of Chemical Reactions

Procedure Overview

- Fill a trough with water, immerse the graduated cylinder to fill it with water, invert the cylinder, and clamp it in place
- Obtain hydrogen peroxide in a clean, dry beaker
- Use the auto-dispenser in the hood to deliver bleach to Erlenmeyer flask
- Place the vial on the balance and press Tare to "zero" the mass reading
- Deliver desired mass of H_2O_2 to the vial (according to the table provided in the procedure) – record the mass exactly as it appears on the balance
- Using tweezers, carefully place the vial upright in the Erlenmeyer flask
- Complete the setup by firmly stoppering the flask and placing the other end of the rubber tubing into the mouth of the inverted graduated cylinder (someone should hold it in place so it doesn't fall out)
- Gently tilt the Erlenmeyer flask to the side, tipping the vial over, and gently swirl the contents of the flask to completely mix
- When bubble formation has stopped, record the volume of water displaced in the graduated cylinder
- Repeat these steps to fill out the table and sketch a plot of "Volume of O_2 " vs. "Mass of H_2O_2 added" in your notebook and make sure you have enough good data values before moving on to part II
- Use the digital pipettor to measure the mass of 500 μL of bleach solution – repeat four times
- Wash the volumetric flask and rinse it well with DI water
- Use the auto-dispenser in the hood to deliver 1, 10-phenanthroline (or phen) solution to the flask – record the concentration of this stock solution
- Add DI water to the phen solution, diluting it to the mark on the neck of the flask
- Transfer the diluted phen solution to a clean, dry beaker
- Use the digital pipettor to deliver 100 μL of Fe^{2+} solution to the phen solution in the beaker
- Mix well and transfer a portion of the solution to the cuvette (note detailed instructions)

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