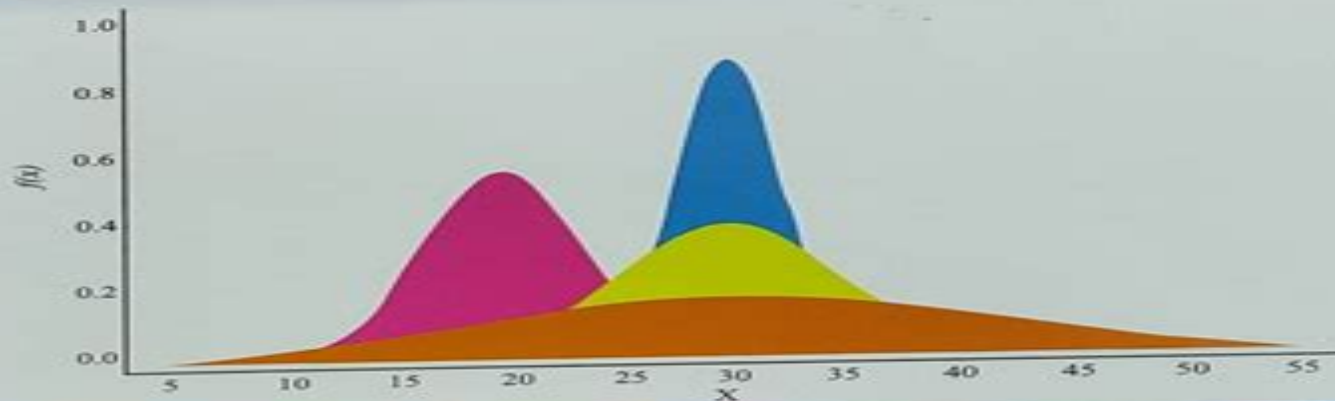


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 $n = \text{number of electrons transferred}$
 $F = 96485 \text{ C/mol}$
 $\Delta G^\circ = -RT \ln K$
 $R = 8.314 \text{ J/(mol}\cdot\text{K)}$
 $T = \text{temperature in Kelvin}$
 $K = \text{equilibrium constant}$
 $E^\circ_{\text{cell}} = \frac{\Delta G^\circ}{-nF}$
 $E^\circ_{\text{cell}} = \frac{RT}{nF} \ln K$
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