

## CHEMISTRY 002

### Question 1:

Using the Maxwell-Boltzmann distribution, derive the expression for the distribution of kinetic energy in an ideal gas. How does the distribution of energy change with temperature and molecular mass? Discuss the physical significance of the temperature dependence in the energy distribution and its connection to specific heat capacities.

### Question 2:

Derive the expressions for the lowering of vapor pressure, elevation of boiling point, depression of freezing point, and osmotic pressure for dilute solutions. Explain how these colligative properties depend on the molarity of the solute and the nature of the solvent. How can these properties be used to determine the molar mass of a non-volatile solute?

### Question 3:

A solution is prepared by dissolving 3.50 g of an unknown non-volatile, non-electrolyte solute in 50.0 g of benzene. The observed freezing point of the solution is  $4.80^{\circ}\text{C}$ , while the freezing point of pure benzene is  $5.50^{\circ}\text{C}$ . The freezing point depression constant for benzene is  $5.12^{\circ}\text{C}\cdot\text{kg/mol}$ . Calculate the molar mass of the solute, and determine the deviation from Raoult's Law for this solution.

### Question 4:

A solution is formed by mixing two non-volatile liquids, A and B, with mole fractions  $x_A$  and  $x_B$ . The vapor pressures of pure A and pure B are 40.0 mmHg and 80.0 mmHg, respectively. If the activity coefficients of A and B in the solution are  $\gamma_A$  and  $\gamma_B$ , respectively, calculate the total vapor pressure of the solution. Additionally, discuss the implications of the non-ideal behavior in terms of positive or negative deviation from Raoult's Law.

### Question 5:

A 1.00 M solution of sodium chloride (NaCl) at  $25^{\circ}\text{C}$  shows an osmotic pressure of 24.5 atm. Calculate the expected osmotic pressure for the same concentration of NaCl in a solution that is

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**Griswold**



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