

CHEM 1211 Fall 2021

Exam 4

- These are the equations that will be provided on exam 4.

Equations

$$w = -P\Delta V$$

$$\Delta E = q + w$$

$$q = mC_p\Delta T$$

$$q_{\text{sur}} = C_{\text{sur}}\Delta T$$

$$\Delta H^\circ = \sum n\Delta H_f^\circ(\text{products}) - \sum n\Delta H_f^\circ(\text{reactants})$$

$$E = h\nu$$

$$c = \lambda\nu$$

$$\Delta E = -2.18 \times 10^{-18} \text{ J} \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$$

$$\lambda = h/mv$$

$$PV = nRT$$

Constants

$$1 \text{ L atm} = 101.325 \text{ J}$$

$$\text{STP} = 273 \text{ K and } 1 \text{ atm}$$

$$R = 0.08206 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} = 8.314 \text{ J}/\text{mol}\cdot\text{K} =$$

$$8.314 \text{ kPa}\cdot\text{L}/\text{mol}\cdot\text{K}$$

$$\text{Avogadro's number: } 6.022 \times 10^{23}$$

$$J = 1 \text{ kg}\cdot\text{m}^2\cdot\text{s}^{-2}$$

$$C_p(\text{water}) = 4.184 \text{ J/g}\cdot^\circ\text{C}$$

$$\text{Mass of an electron: } 9.109 \times 10^{-31} \text{ kg}$$

$$\text{Mass of a proton: } 1.673 \times 10^{-27} \text{ kg}$$

$$c = 3.00 \times 10^8 \text{ m/s}$$

$$h = 6.626 \times 10^{-34} \text{ Js}$$

$$R_H = 2.18 \times 10^{-18} \text{ J}$$

Temperature

$$0 \text{ K} = -273.15^\circ\text{C}$$

$$\text{K} = ^\circ\text{C} + 273.15$$

$$^\circ\text{C} = (5/9)(^\circ\text{F} - 32)$$

Volume

$$1 \text{ mL} = 1 \text{ cm}^3$$

Pressure

$$1 \text{ pascal} = 1 \text{ N/m}^2 = 1 \text{ kg}/\text{m}\cdot\text{s}^2$$

$$1 \text{ atm} = 101.325 \text{ kPa} = 760 \text{ mm Hg} = 760$$

$$\text{torr} = 14.70 \text{ lb}/\text{in}^2$$

$$1 \text{ bar} = 1 \times 10^5 \text{ Pa (exactly)}$$

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