

Quantity	Symbol	Value	Unit	Relative uncertainty (ppm)
PHYSICO-CHEMICAL CONSTANTS				
Avogadro constant	N_A, L	6.022 136 7(36)	10^{23} mol^{-1}	0.59
atomic mass constant	m_a	1.660 540 2(10)	10^{-27} kg	0.59
$m_a = \frac{1}{12} m(^{12}\text{C})$		931.494 32(28)	MeV	0.30
in electron volts: $m_a c^2 / \{e\}$	F	96 485.309(29)	C mol $^{-1}$	0.30
Faraday constant $N_A e$	$N_A h$	3.990 313 23(36)	$10^{-10} \text{ J s mol}^{-1}$	0.089
molar Planck constant	$N_A h c$	0.119 626 58(11)	J m mol $^{-1}$	0.089
molar gas constant	R	8.314 510(70)	J mol $^{-1} \text{ K}^{-1}$	8.4
Boltzmann constant R/N_A	k	1.380 658(12)	$10^{-23} \text{ J K}^{-1}$	8.5
in electron volts: $k/\{e\}$		8.617 385(73)	$10^{-5} \text{ eV K}^{-1}$	8.4
in hertz: k/h		2.083 674(18)	$10^{10} \text{ Hz K}^{-1}$	8.4
in wavenumbers: k/hc		69.503 87(59)	m $^{-1} \text{ K}^{-1}$	8.4
molar volume (ideal gas) RT/p	V_m	0.022 414 10(19)	m $^3 \text{ mol}^{-1}$	8.4
$T = 273.15 \text{ K}, p = 101\,325 \text{ Pa}$		2.686 763(23)	10^{26} m^{-3}	8.5
Loschmidt constant N_A/V_m	V_m	0.022 711 08(19)	m $^3 \text{ mol}^{-1}$	8.4
$T = 273.15 \text{ K}, p = 100 \text{ kPa}$				
Sackur-Tetrode constant	S_0/R	-1.151 693(21)		18
(absolute entropy constant) ^b		-1.164 856(21)		18
$\frac{5}{2} + \ln[(2\pi m_a k T_1/h^2)^{3/2} k T_1/p_0]$				
$T_1 = 1 \text{ K}, p_0 = 100 \text{ kPa}$				
$T_1 = 1 \text{ K}, p_0 = 101\,325 \text{ Pa}$				
Stefan-Boltzmann constant	σ	5.670 51(19)	$10^{-8} \text{ W m}^{-2} \text{ K}^{-4}$	34
$(\pi^2/60)k^4/h^3c^2$	c_1	3.741 774 9(22)	10^{-16} W m^2	0.60
first radiation constant $2\pi\hbar c^2$	c_2	0.014 387 69(12)	m K	8.4
second radiation constant hc/k				
Wien displacement law constant	b	2.897 756(24)	10^{-3} m K	8.4
$b = \lambda_{\text{max}} T = c_2/4.965\,114\,23\dots$				

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WJ Hussar



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Physicochemical Properties of Ionic Liquid Mixtures Suojiang Zhang, Qing Zhou, Xingmei Lu, Yuting Song, Xinxin Wang, 2017-10-09 This handbook is a valuable resource for scientists engineers graduate students managers decision makers and those who are interested in ionic liquids Many industrial applications rely on the use of Ionic Liquid Mixtures as in solar energy storage waste recycling or batteries Physicochemical Properties of Ionic Liquid Mixtures is a useful handbook that contains the following features the physicochemical properties and property models of mixtures containing ionic liquids supplemented by a comprehensive database of properties listing ionic liquid systems collected from more than 800 dependable literature sources over 60 000 data entries on 39 types of physicochemical properties for 1388 mixtures including binary ternary quaternary and other mixtures *The Journal of Physical Chemistry: Volume 67, Number 1 January 22, 1963* W. Albert Noyes, Jr., 1963 **Journal of the Society of Chemical Industry** Society of Chemical Industry (Great Britain), 1887 **Dechema Chemistry Data Series**, 1987 *The Journal of Physical Chemistry*, 1898 **Recommended Data of Selected Compounds and Binary Mixtures** Karl Stephan, 1987 **Physico-chemical Methods** Joseph Reilly, William Norman Rae, 1954 Molecular Volumes in Chemistry and Biology John C. McGowan, Alan Mellors, 1986 Chemical Engineering: Chemical engineering design John Metcalfe Coulson, John Francis Richardson, 1993 **Acoustic and Thermodynamic Properties of Simple Organic Liquids at Pressures Up to 3 Kbar** Sun Tongfan, 1989 **Indian Journal of Chemistry**, 1997 **Perry's Chemical Engineers' Handbook** Robert H. Perry, Don W. Green, James O.

Maloney,1984 From the fundamentals to details on computer applications and control this handbook provides unrivaled state of the art coverage of all aspects of chemical engineering The seventh edition is completely updated and includes new topics such as biochemical engineering waste management plant safety analysis of plant performance and handling of hazardous materials Over 1 700 illus Copyright Libri GmbH All rights reserved **Physical Methods of Organic Chemistry** ,1959

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