

2010 Brains Physical Chemistry Review-Problem Set 4

Solve the following problems:

- At a given time, N_2 is reacting with H_2 at a rate of 1.50 mol/min to produce NH_3 . At the same time,
 - What is the rate at which the other reactant is changing, and
 - the rate at which the product is changing?



- Initial Rates Method:** Consider a hypothetical reaction, $X + Y \longrightarrow XY$ studied at 20°C .

Concentration (mol L^{-1})		Rate of Disappearance of Y ($\text{mol L}^{-1} \text{ s}^{-1}$)
[X]	[Y]	
1.34×10^{-1}	4.00	1.882
2.69×10^{-1}	4.00	3.760
2.69×10^{-1}	1.20×10^{-1}	11.280

Write the rate law expression for the reaction. Include the value of k and its correct units.

- Dinitrogen tetroxide decomposes spontaneously at 303K in the gas phase:



The rate law expression is $\text{rate} = k[N_2O_4]$, where k is the rate constant $= 5.1 \times 10^{-6} \text{ s}^{-1}$ at 303K . Energy of activation, $E_a = 54.0 \text{ kJ/mol}$

- What is the half-life, $t_{1/2}$ of the N_2O_4 decomposition?
 - Calculate the time (in seconds) required for the partial pressure of $N_2O_4(g)$ to decrease from 0.10 atm to 0.020 atm at $T=303 \text{ K}$.
 - If the E_a does not vary considerably between 303K and 340K , calculate the new rate constant, k_2 , at 340K .
- A reaction given below follows a rate law of $\text{rate} = k[A]^2$

$$A \longrightarrow \text{decomposed products.} \quad T = 300\text{K}$$

Half-life, $t_{1/2} = 80 \text{ s}$
Initial concentration of $A = 5.0 \text{ M}$

 - What is the rate constant at 300 K ?
 - How much of A has reacted after 200 s (at $T = 300 \text{ K}$)?
 - How long will it take until only one-fourth of the initial concentration of A is left?
 - How much faster/slower is the reaction at 500 K if $E_a = 100 \text{ kJ/mol}$.

- The thermal decomposition of ammonia at high temperature was studied in the presence of inert gases. The data given below are for a single experiment:



Time (hour)	NH_3 (mol/L)
0	8.00×10^{-2}
25	6.75×10^{-2}
50	5.84×10^{-2}
75	5.15×10^{-2}

A plot of the appropriate concentration expression against time indicate that the decomposition is a **second-order** reaction, $\text{rate} = k[NH_3]^2$

- Find the rate constant, k , of the reaction from the slope of the line.

Physical Chemistry Prob Sol

**Victor E Borisenko, Sergei Vasil'evich
Gaponenko, Valerij S Gurin**



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