

# Chem 132

## Day 1 - Ch. 13

- **Unsaturated solution** - not as much solute dissolved as possible (below the limit)
  - **Saturated Solution** - maximum amount of solute dissolved (at the limit)
  - **Supersaturated solution** - more than the maximum amount of solute (over the limit)
  - **Dilute solution** - little solute and lots of solvent
  - **Concentrated solution** - lots of solute in the given solvent
  - **Concentration** = amount of solute / amount of solvent or solution
    - Solution = solute + solvent
    - Aq = solute + H<sub>2</sub>O
  - **Mass Percent** = (mass of solute / mass of solution) × 100
  - 8g NaOH are dissolved in 50g water. Calculate the mass percent.
    - $(8\text{g NaOH} / (8\text{g} + 50\text{g})) \times 100 = 13.8\%$
  - What weight of KCl and water are needed to make 250g of a 5% KCl solution?
    - $5 = (\text{g solute} / 250 \text{ g solution}) \times 100 = 12.5\text{g KCl}$
    - $250 - 12.5 = 237.5 \text{ water}$
  - Calculate the mole fraction of NH<sub>3</sub> for a solution of 30.0g of NH<sub>3</sub> in 70.0g H<sub>2</sub>O.
    - **Mole fraction** = mol A / total mol
    - $30 \text{ g NH}_3 / 17\text{g} = 1.76 \text{ mol NH}_3$
    - $70 \text{ g water} / 18\text{g} = 3.89 \text{ mol water}$
    - $1.76 / (1.76 + 3.89) = 0.312$
  - **Molarity** - M - moles of solute / Liters of solution
  - **Molality** - m - moles of solute / kg of solvent
  -
- H<sub>2</sub>O is always the solvent!
- You have a solution that is 20% by mass MgSO<sub>4</sub> and the density of the solution is 1.30g/mL.
    - **Density** = mass / volume
    - $20\text{g MgSO}_4 / 120\text{g MgSO}_4 = 0.167 \text{ mol MgSO}_4$
    - $80\text{g} = 0.08 \text{ kg water}$
    - $\text{Molality} = 0.167 \text{ mol} / 0.08\text{kg} = 2.09 \text{ m}$
    - $80\text{g water} / 18\text{g} = 4.44 \text{ mol H}_2\text{O}$
    - $\text{Mole Fraction} = 0.167 / (0.167 + 4.44) = 0.0360 \text{ MgSO}_4$
    - $100\text{g} / 1.30\text{g} / 1000 = 0.0769\text{L}$
    - $\text{Molarity} = 0.167 / 0.0769 \text{ L} = 2.17 \text{ M}$
  - The density of a 1.099 M HCHO<sub>2</sub> solution is 1.0115g/mL.
    - $1.099 \text{ mol HCHO}_2 \times 46\text{g} = 50.6\text{g HCHO}_2$
    - $1 \text{ L} \times 1000 \text{ mL} \times 1.0115\text{g} = 1011.5\text{g solution}$
    - $\text{Mass \%} = (50.6\text{g} / 1011.5\text{g}) \times 100 = 5\%$
    - $\text{Molality} = 1.099 / 0.9609 \text{ kg} = 1.14 \text{ m}$

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