

# CHE-112 Practice Exam 4 Answers at End

1. In which one of these pairs will the entropy of the first substance be greater than that of the second? Assume  $P$  and  $T$  are the same for each pair, unless stated otherwise.

- A) 1 mole of  $\text{F}_2(g)$ ; 1 mole of  $\text{Cl}_2(g)$
- B) 1 mole of  $\text{I}_2(s)$ ; 1 mole of  $\text{I}_2(g)$
- C) 1 mole of  $\text{CaCO}_3(s)$ ; 1 mole of  $\text{CaO}(s)$  plus 1 mole of  $\text{CO}_2(g)$
- D) 1 mole of  $\text{H}_2(g)$  at  $25^\circ\text{C}$ ; 1 mole of  $\text{H}_2(g)$  at  $50^\circ\text{C}$
- E) 1 mole of  $\text{O}_2(g)$ ; 1 mole of  $\text{O}_3(g)$

2. For a chemical reaction to be spontaneous only at high temperatures, which of the following conditions must be met?

- A)  $\Delta S^\circ > 0$ ,  $\Delta H^\circ > 0$
- B)  $\Delta S^\circ > 0$ ,  $\Delta H^\circ < 0$
- C)  $\Delta S^\circ < 0$ ,  $\Delta H^\circ < 0$
- D)  $\Delta S^\circ < 0$ ,  $\Delta H^\circ > 0$
- E)  $\Delta G^\circ > 0$

3. Which of the following is necessary for a process to be spontaneous?

- A)  $\Delta H_{\text{sys}} < 0$
- B)  $\Delta S_{\text{sys}} > 0$
- C)  $\Delta S_{\text{univ}} < 0$
- D)  $\Delta S_{\text{univ}} > 0$
- E)  $\Delta G_{\text{sys}} = 0$

4. Which of the following is always true for an endothermic process?

- A)  $q_{\text{sys}} > 0$ ,  $\Delta S_{\text{univ}} < 0$
- B)  $q_{\text{sys}} < 0$ ,  $\Delta S_{\text{univ}} > 0$
- C)  $q_{\text{sys}} < 0$ ,  $\Delta S_{\text{univ}} < 0$
- D)  $q_{\text{sys}} > 0$ ,  $\Delta S_{\text{univ}} > 0$
- E)  $w < 0$

5. Which relationship or statement best describes  $\Delta S^\circ$  for the following reaction?



- A)  $\Delta S^\circ \approx 0$
- B)  $\Delta S^\circ < 0$
- C)  $\Delta S^\circ > 0$
- D)  $\Delta S^\circ = \Delta H^\circ/T$
- E) More information is needed to make a reasonable prediction.

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