

Final Exam (Practice)

Chemistry 111, Prof. Metz

120 minutes

Put your name and student ID# on both pages of the answer sheet.

Put all answers on the answer sheet - answers for questions 30-32 go on page 2.

Turn in the entire test.

For each question, only one answer is correct.

Useful information is on page 2 of the answer sheet.

Each multiple choice question is worth 3 points.

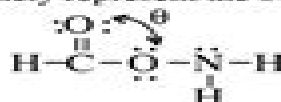
1. The protein randomase has a molecular mass (molecular weight) of 1.6×10^5 . One microgram of this protein costs \$100. How many molecules of randomase are in one microgram of randomase?

- A) 1.60×10^{11} B) 3.76×10^{12} C) 6.25×10^{12} D) 6.02×10^{17} E) 3.76×10^{18}

2. The percentage composition of sulfur in sodium sulfate is

- A) 22.6 % B) 25.4 % C) 26.9 % D) 29.8 % E) 41.1 %

3. Consider the compound whose Lewis dot structure is drawn below (note that this is merely the Lewis dot structure - it does NOT accurately represent the bond angles in the molecule).



The O-C-O bond angle (eg, the bond angle about the carbon atom) is labelled θ . This angle is approximately

- A) 45° B) 90° C) 109° D) 120° E) 180°

4. In an experiment similar to the one you did in Lab, 14.00 grams of molybdenum metal react with sulfur, producing 21.02 g of a compound containing only molybdenum and sulfur. What compound is produced?

- A) Mo_2S B) MoS C) Mo_2S_3 D) MoS_2 E) MoS_3

5. Photodissociation of NO_2 (a major component of smog) in cities eventually leads to formation of irritants.

The photodissociation



requires photons with a wavelength of 398 nm. This wavelength corresponds to an energy of _____ kJ/mole. (Hint: Watch the units! Find the energy in ONE photon, then calculate the energy in a MOLE of photons)

- A) 5×10^{-19} B) 5×10^{-16} C) 5×10^{-4} D) 0.3 E) 300

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