

**CHEM 1411 – General Chemistry I**  
**Practice Problems, Chapters 1–3**

**Chapter 1 – Chemistry: The Study of Change**

1. Element, compound, homogeneous mixture (solution), or heterogeneous mixture:
  - a) orange juice  
(freshly-squeezed)
  - b) brass
  - c) 0.9% saline (NaCl) solution
  - d) garden soil
  - e) room air
  - f) methane gas
  - g) sodium metal
  - h) N<sub>2</sub> gas
  - i) Cu(NO<sub>3</sub>)<sub>2</sub> crystals
  
2. Define (some of these terms are found in Chapters 2 and 3):
  - a) element
  - b) compound
  - c) solution
  - d) physical change
  - e) chemical change
  - f) law of conservation of mass
  - g) intensive properties
  - h) extensive properties
  - i) precision
  - j) accuracy
  - k) isotopes
  - l) stoichiometry
  - m) the mole
  - n) Avogadro's number
  - o) limiting reactant
  
3. Convert
  - a) 5 feet 10 inches to meters
  - b) 55 miles per hour to cm per second
  - c) 5.0 quarts to liters
  - d) 1.00 ft<sup>2</sup> to cm<sup>2</sup>
  - e) 32.06 amu to g
  - f) 25 g of carbon to moles of carbon
  - g) 1.00 g of gold to number of atoms of gold
  - h) 25 °F to Kelvins
  
4. How many significant figures are in the following numbers?
  - a) 3.1416
  - b)  $6.022 \times 10^{23}$  atoms/mole
  - c) 0.000104 g
  - d) 6.0 g/mL
  - e) 1.00794 amu
  - f) 5000 mi

# Pkgacp Cer General Chemistry 1 Chem 1411

**M Mosston**



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