

1. Ethanol reacts with H^+ and H_2O in a general equilibrium as follows (shown below):

- a. 0
- b. 1
- c. 100%
- d. 10%

2. Ethanol reacts with H^+ and H_2O in a general equilibrium as follows (shown below):

- a. 0%
- b. 10%
- c. 100%
- d. 10%

3. In general, the equilibrium constant K for a reaction is defined as the ratio of the products to the reactants, each raised to the power of its stoichiometric coefficient.

- a. $K = \frac{[\text{products}]}{[\text{reactants}]}$
- b. $K = \frac{[\text{reactants}]}{[\text{products}]}$
- c. $K = \frac{[\text{products}]}{[\text{reactants}]}$
- d. $K = \frac{[\text{reactants}]}{[\text{products}]}$

4. The equilibrium constant K for a reaction is defined as the ratio of the products to the reactants, each raised to the power of its stoichiometric coefficient.

- a. $K = \frac{[\text{products}]}{[\text{reactants}]}$
- b. $K = \frac{[\text{reactants}]}{[\text{products}]}$
- c. $K = \frac{[\text{products}]}{[\text{reactants}]}$
- d. $K = \frac{[\text{reactants}]}{[\text{products}]}$

5. The equilibrium constant K for a reaction is defined as the ratio of the products to the reactants, each raised to the power of its stoichiometric coefficient.

- a. $K = \frac{[\text{products}]}{[\text{reactants}]}$
- b. $K = \frac{[\text{reactants}]}{[\text{products}]}$
- c. $K = \frac{[\text{products}]}{[\text{reactants}]}$
- d. $K = \frac{[\text{reactants}]}{[\text{products}]}$

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