

(c) A series of equilibrium mixtures is prepared in order to calculate K_{eq} . Absorbance values were measured with the spectrophotometer and recorded in the table.

	E1	E2	E3	E4
Volume of $2.00 \times 10^{-3} \text{ M}$ $\text{Fe}(\text{NO}_3)_3$ (mL)	25.00	25.00	25.00	25.00
Volume of $2.00 \times 10^{-3} \text{ M}$ NaSCN (mL)	0	5.00	15.00	25.00
Total solution volume (mL)	50.00	50.00	50.00	50.00
Absorbance of equilibrium mixture	0	0.137	0.365	0.488
$[\text{FeNCS}^{2+}]$ (M) (Use your plot to determine this)				

(d) Calculate the concentrations and K_{eq} for the ICE table of E2.

	$\text{Fe}^{3+}(\text{aq})$	$+$	$\text{SCN}^{-}(\text{aq})$	\rightleftharpoons	$\text{FeNCS}^{2+}(\text{aq})$
E2 Initial, M	_____		_____		_____
Change, M	_____		_____		_____
Equilibrium, M	_____		_____		_____

$$K_{eq} = \underline{\hspace{2cm}}$$

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