

- 11.59 Compare the strengths of the following pairs of acids: (a) H_2SO_4 and H_2SeO_4 , and (b) H_3PO_4 and H_3AsO_4 .
- 11.60 Which of the following is the stronger acid, CH_2ClCOOH or CHCl_2COOH ? Explain your choice.
- 11.61 Consider phenol and methanol (both of which are alcohols):



Experimentally, phenol is found to be a stronger acid than methanol. Explain this difference in terms of the structures of the conjugate bases. (*Hint:* A more stable conjugate base favors ionization. Only one of these conjugate bases can be stabilized by resonance.)

Many Salts Have Acid-Base Properties in Aqueous Solution

- 11.62 Predict the pH (>7 , $=7$, or <7) of aqueous solutions containing the following salts: (a) KBr , (b) $\text{Al}(\text{NO}_3)_3$, (c) BaCl_2 , and (d) $\text{Bi}(\text{NO}_3)_3$.
- 11.63 Predict whether the following aqueous solutions are acidic, basic, or nearly neutral: (a) NaBr , (b) $\text{K}_2\text{S}_2\text{O}_8$, (c) NH_4NO_3 , and (d) $\text{Cr}(\text{NO}_3)_3$.
- 11.64 A certain salt, MX (containing the M^+ and X^- ions), is dissolved in water, and the pH of the resulting solution is 7.0 at 25°C . Can you say anything about the strengths of the acid and the base from which the salt is derived?
- 11.65 In a certain experiment, a student finds that the pH values of 0.10 M solutions of three potassium salts KX , KY , and KZ are 7.0, 9.0, and 11.0, respectively. Arrange the acids HX , HY , and HZ in the order of increasing acid strength.
- 11.66 Calculate the pH of a 0.36 M CH_3COONa solution at 25°C .
- 11.67 Calculate the pH of a 0.42 M NH_4Cl solution at 25°C .
- 11.68 Predict the pH (>7 , $=7$, <7) of an NaHCO_3 solution.
- 11.69 Predict whether a solution containing the salt K_2HPO_4 will be acidic, basic, or neutral.
- 11.70 Calculate the pH of a 0.015 M oxalic acid solution at 25°C .

Oxide and Hydroxide Compounds Can Be Acidic or Basic in Aqueous Solution Depending on Their Composition

- 11.71 Arrange the oxides in each of the following groups in order of increasing basicity: (a) K_2O , Al_2O_3 , BaO , and (b) CrO_3 , CrO , Cr_2O_3 .
- 11.72 $\text{Zn}(\text{OH})_2$ is an amphoteric hydroxide. Write balanced ionic equations to show its reaction with (a) HCl , and (b) NaOH [the product is $\text{Zn}(\text{OH})_4^{2-}$].

- 11.73 $\text{Al}(\text{OH})_3$ is insoluble in pure water, but it dissolves in excess NaOH in solution. Write a balanced ionic equation for this reaction. What type of reaction is this?

Additional Problems

- 11.74 The K_a of formic acid is 1.7×10^{-4} at 25°C . Will the acid become stronger or weaker at 40°C ? Explain.
- 11.75 A typical reaction between an antacid and the hydrochloric acid in gastric juice is
- $$\text{NaHCO}_3(s) + \text{HCl}(aq) \rightleftharpoons \text{NaCl}(aq) + \text{H}_2\text{O}(l) + \text{CO}_2(g).$$
- Calculate the volume (in L) of CO_2 generated from 0.350 g of NaHCO_3 and excess gastric juice at 1.00 bar and 37.0°C .
- 11.76 To which of the following would the addition of an equal volume of 0.60 M NaOH lead to a solution having a lower pH: (a) water, (b) 0.30 M HCl , (c) 0.70 M KOH , and (d) 0.40 M NaNO_3 ?
- 11.77 The pH of a 0.0642 M solution of a monoprotic acid is 3.86 at 25°C . Is this a strong acid?
- 11.78 Like water, liquid ammonia undergoes autoionization:



- (a) Identify the Brønsted acids and Brønsted bases in this reaction. (b) What species correspond to H^+ and OH^- , and what is the condition for a neutral solution?
- 11.79 HA and HB are both weak acids, although HB is the stronger of the two. Will it take a larger volume of a 0.10 M NaOH solution to neutralize 50.0 mL of 0.10 M HB or 50.0 mL of 0.10 M HA ?
- 11.80 A solution contains a weak monoprotic acid HA and its sodium salt NaA , both at 0.1 M concentrations. Show that $[\text{OH}^-] = K_a/K_b$.
- 11.81 The three common chromium oxides are CrO , Cr_2O_3 , and CrO_3 . If Cr_2O_3 is amphoteric, what can you say about the acid-base properties of CrO and CrO_3 ?
- 11.82 Use the data in Appendix 2 to determine ΔH° , ΔS° , and ΔG° for the autoionization of water at 25°C . Using the techniques discussed in Chapter 10, determine the value of K_w at 25°C and at 60°C . What assumptions must you make to do these calculations? Compare your results with the data in Table 11.2.
- 11.83 Use the data in Table 11.3 to calculate the equilibrium constant for the following reaction at 25°C :
- $$\text{HCOOH}(aq) + \text{OH}^-(aq) \rightleftharpoons \text{HCOO}^-(aq) + \text{H}_2\text{O}(l)$$
- 11.84 Use the data in Table 11.3 to calculate the equilibrium constant for the following reaction at 25°C :
- $$\text{CH}_3\text{COOH}(aq) + \text{NO}_3^-(aq) \rightleftharpoons \text{CH}_3\text{COO}^-(aq) + \text{HNO}_3(aq)$$

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