

d. Balancing oxidation-reduction reactions.

Half Reaction

losing
 e^{-}

Electron Transfer

X is transferring e^{-}
to Ygaining
 e^{-} Y is accepting e^{-}
from X

Result

X is oxidized
(reducing agent)Y is reduced
(oxidizing agent)

e. If this still doesn't make sense (and don't worry – you'll need to study it for a bit before it often clicks!) try this definition:

oxidation = increase in oxidation numberreduction = decrease in oxidation number

Types of Oxidation-Reduction Reactions (through examples)

1. Combination reaction:

<ox> \uparrow O.N.<red> \downarrow O.N.Reducing Agent: Al
Oxidizing agent: Br₂

2. Decomposition reaction:



$$0 = (+1) + x + 3(-2)$$

$$x = +5$$



r.a.

o.a.

o.a. Cl

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
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