

Chemistry 1100 Practice FINAL EXAM KEY

The data below may be useful. Periodic Tables are available.

Gas constant, $R = 0.082057 \text{ L atm K}^{-1} \text{ mol}^{-1}$ $1 \text{ atm} = 760 \text{ torr}$
 $= 8.3144 \text{ J mol}^{-1} \text{ K}^{-1} = 1.987 \text{ cal mol}^{-1} \text{ K}^{-1}$.

Graham's law: rate of effusion $\propto \frac{1}{\sqrt{M}}$

Molal freezing point depression constant, K_f , for water = -1.86°C/m .

The freezing point depression is given by $\Delta T = i m K_f$

where m (*molarity*) is the number of moles of involatile, non-dissociating solute per kg of solvent.

$1 \text{ lb} = 454 \text{ g}$ $1 \text{ ton} = 2000 \text{ lb}$ $1 \text{ qt} = 946 \text{ mL}$ $1 \text{ meter} = 1.0936 \text{ yard}$

$1 \text{ joule} = 1 \text{ kg}\cdot\text{m}^2 \cdot \text{s}^{-2}$ Planck's constant = $6.626068 \times 10^{-34} \text{ m}^2 \text{ kg} / \text{s}$

Speed of light, $c = 3.0 \times 10^{10} \text{ cm s}^{-1} = 3.00 \times 10^8 \text{ m/s}$ $E = h\nu$

Planck's constant, $h = 6.626 \times 10^{-34} \text{ J s} = 6.626 \times 10^{-34} \text{ m}^2 \text{ kg} / \text{s}$

$m_e = 9.11 \times 10^{-31} \text{ kg}$ $c = \lambda\nu$ $p\lambda = h$

$$\Delta E = Z^2 R_H \cdot \left[\frac{1}{n_i^2} - \frac{1}{n_f^2} \right]; \quad R_H = 2.179 \times 10^{-18} \text{ J}$$

$$\frac{1}{\lambda} = Z^2 R \cdot \left[\frac{1}{n_i^2} - \frac{1}{n_f^2} \right]; \quad R = 1.097 \times 10^7 \text{ m}^{-1} \quad Z=1 \text{ for H atom.}$$

Raoult's Law says $p_i = X_i P_{\text{total}}$

Henry's law: $C_{\text{gas}} = k P_{\text{gas}}$

Osmotic Pressure $\pi V = nRT$ or $\pi = MRT$

$1 \text{ \AA} = 10^{-10} \text{ m}$

Distance to nearest star = $4 \times 10^{26} \text{ \AA}$

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