

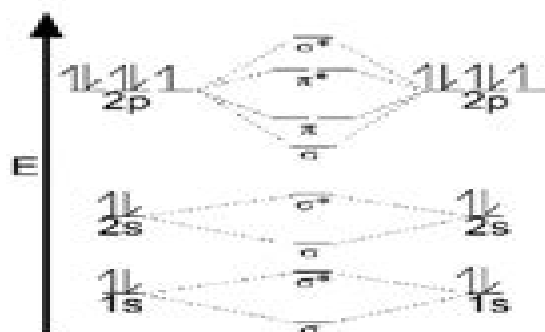
CEM 141 Mock Exam 3 Questions – FS24

This mock exam is intended to give you an idea of some of the different types of questions that could be asked on an exam. It does not cover all of the material that may appear on your exam.

Exam 3 will cover all of the following learning objectives: 3.5, 3.7–3.13, 4.1–4.11, and 4.16–4.17.

Multiple Choice Questions

- At room temperature, substance A is a solid and substance B is a gas. A and B react according to the following equation: $A + 3B \rightarrow C$
At room temperature, what state of matter do you predict for substance C?
 - Solid, because C is heavier than A or B.
 - Gas, because C is three parts B (which is a gas) and one part A (which is a solid).
 - Liquid, because C will be intermediate between a solid and a gas.
 - We cannot predict the state of matter because we don't know about the bonding/interactions present within C.
- To the right is an incomplete molecular orbital diagram for a diatomic molecule. The electrons are shown in the atomic orbitals, but not within the molecular orbitals. Add the electrons to the molecular orbitals. How many bonds exist between the atoms in this diatomic molecule?
 - 0 bonds
 - 1 bond
 - 2 bonds
 - 3 bonds
 - 4 bonds
 - 5 bonds
- What is the best explanation for why metals are ductile?
 - Valence electrons carry charge anywhere within molecular orbitals that span the metal.
 - Photons are absorbed and re-emitted.
 - Metallic bonds involve delocalized electrons whereas covalent bonds involve localized electrons.
 - When the nuclei move relative to one another, the valence electrons act as a glue holding them together.



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