

# CHEM 1211 Exam 5

## Multiple Choice

Identify the choice that best completes the statement or answers the question.

Questions 1-30 are worth 6.67 points each for a total of 200 points.

1. Calcium hydroxide will neutralize phosphoric acid to produce solid calcium phosphate and water. Write and balance the chemical equation for this reaction. What is the sum of the coefficients?

- a. 4
- b. 11
- ☒ c. 12
- d. 16
- e. 28



Nitrogen gas reacts with hydrogen gas to produce gaseous ammonia:



Use the chemical equation to answer questions 2 through 5. The molecular weights of  $\text{N}_2$ ,  $\text{H}_2$ , and  $\text{NH}_3$  are 28.014 g/mol, 2.016 g/mol, and 17.031 g/mol.

2. How many moles of ammonia could be formed from the reaction of 3.25 moles of hydrogen gas with excess nitrogen gas?

- a. 1.63 moles
- ☒ b. 2.17 moles
- c. 3.25 moles
- d. 4.88 moles
- e. 6.50 moles

$$3.25 \text{ mol H}_2 \cdot \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} =$$

B

3. How many grams of nitrogen gas would be needed to completely react 4.52 grams of hydrogen gas?

- a. 0.976 grams
- b. 4.52 grams
- ☒ c. 20.9 grams
- d. 62.8 grams
- e. 188 grams

$$4.52 \text{ g H}_2 \cdot \frac{1 \text{ mol H}_2}{2.016 \text{ g H}_2} \cdot \frac{1 \text{ mol N}_2}{3 \text{ mol H}_2} \cdot \frac{28.014 \text{ g}}{1 \text{ mol N}_2}$$

C

4. How many grams of ammonia could be produced from the reaction of 5.23 grams of nitrogen with 5.23 grams of hydrogen?

- a. 0.445 grams
- b. 3.79 grams
- ☒ c. 7.58 grams
- d. 35.1 grams

$$5.23 \text{ g H}_2 \cdot \frac{1 \text{ mol H}_2}{2.016 \text{ g H}_2} \cdot \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \cdot \frac{17.031 \text{ g}}{1 \text{ mol}}$$

C

e) 52.6 grams

$$5.23 \text{ g N}_2 \cdot \frac{1 \text{ mol N}_2}{28.014 \text{ g N}_2} \cdot \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \cdot \frac{17.031 \text{ g}}{1 \text{ mol}}$$

# Pkgacp Chem 1211

**Clarence Walker Barron**



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