

NAME: _____

Instructions: Answer all questions. You should edit this DOCX file and insert your answers herein (either by typing, or by scanning/inserting handwritten work). You may use Excel or similar programs to generate charts/graphs/Tables, and similarly, insert your work into the respective portion of the exam. When you have completed your exam, please convert it to a PDF file. The exam is due by email to michael.ketterer@nau.edu by 11:59 PM on Thursday, September 22.

1. (12 points) A gold crown has the following mass composition:

Au	78.5 %	Ag	18.3 %
Cu	1.5 %	Pd	1.2 %
Pt	0.5 %	Others	< 0.1 %

A 0.649 g sample of this material is dissolved using HCl and HNO₃ and is diluted in a Class A volumetric flask to a volume of 100.0 ± 0.1 mL.

Calculate the Ag concentration in milligrams/liter (mg/L) in the resulting solution.

**Concentration of Ag in mg/l = $(0.12 \times 1000)/100 = 1.2$ gm/lt
= 1.2×10^3 mg/lt. (1 gm = 1000 mg).**

Calculate the molarity (M) of Au in the resulting solution. Atomic mass of Au = 197.0 g/mole

Amount of Au is 78.5% that is $(0.649 \times 78.5)/100$

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