Gas

| Kinetic Theory of Gas

Ulmit	Definition	
N/m^2	$P = \frac{F}{A}$	
27	$X_a = \frac{n_a}{n_{total}}$	
N/m^2	$P_a = X_a P_{\text{total}}$	
J	$W = -P\Delta V$	
*:	$z = \frac{PV}{nRT}$	
	N/m ² N/m ² J	N/m^2 $P = \frac{F}{A}$ $X_a = \frac{n_a}{n_{total}}$ N/m^2 $P_a = X_a P_{total}$ J $W = -P\Delta V$

Description	Equations
Ideal gas law	PV = nRT
Ideal gas constant and Boltzmann constant	$R = N_A k_B$
Average kinetic energy of one gas molecule	$\varepsilon = \frac{3}{2}k_BT$
Root mean square speed gas	$v_{\rm RMS} = \sqrt{\frac{3RT}{\mathcal{M}}} = \sqrt{\frac{3k_BT}{m}}$
Mean speed of gas	$v = \sqrt{\frac{8RT}{\pi M}} = \sqrt{\frac{8k_BT}{\pi m}}$
Most probable speed of gas	$v_{\text{exp}} = \sqrt{\frac{2RT}{M}} = \sqrt{\frac{2k_BT}{m}}$
Maxwell-Boltzmann speed distribution	$f(v) = 4\pi \left(\frac{m}{2\pi k_B T}\right) v^2 \exp\left(\frac{-mv^2}{2k_B T}\right)$
Van der Waals equation of state	$(P + a\frac{n^2}{V^2})(V - bn) = nRT$
Lennard-Jones potential	$V_{LJ} = 4\varepsilon \left(\left(\frac{\sigma}{R} \right)^{12} - \left(\frac{\sigma}{R} \right)^{6} \right)$

| Molecular Collisions and Rate Processes

Description	Equations
Molecule collision rate with wall	$Z_w \propto \frac{1}{4} \frac{N}{V} v A = \frac{1}{4} \frac{N}{V} \sqrt{\frac{8RT}{\pi M}} A$
Graham's law of effusion	$\frac{\text{rate of effusion of A}}{\text{rate of effusion of B}} = \frac{N_A}{N_B} \sqrt{\frac{M_B}{M_A}}$
Molecule-molecule collision rate	$Z_1 = 4 \frac{N}{V} d^2 \sqrt{\frac{\pi R T}{A A}}$
Mean free path	$\lambda = \frac{\varrho}{Z_1} = \frac{V}{\sqrt{2\pi d^2 N}}$

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