

You will have 110 minutes for this test. All answers are to be submitted on the test paper, which has 4 pages. Scrap paper will be provided, but will not be marked. A periodic table and data tables will be provided. Simple scientific calculators are permitted, but GRAPHING CALCULATORS ARE NOT PERMITTED. Cell phones, smart watches, headphones, translators and other electronic devices are not permitted. These must be turned off and stored in a bag or purse under your desk.

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**PART 1: [23 Marks]** Multiple choice, circle the letter beside the correct response or fill in the correct answer on the blank. In this section do not consider significant figures unless specifically directed to in the question. One mark each, unless otherwise specified.

1. There is 100 mL a stock solution of 0.499 M NaOH available and you need 50.0 mL of 0.127 M solution. How many mL of stock solution would be required? 12.7 mL

2. What is the formal charge of Fe in  $\text{Fe}_2\text{O}_3$ ? 3+

3. How many neutrons in  $^{183}\text{Pr}$ ? 117 neutrons

4. a) Draw the best Lewis structure of  $\text{COCl}_2$  (C central) in the space to the left and complete the table. (6 marks)

Lewis structure: total valence  $e^- = 4 + 6 + 14 = 24$



a) molecular geometry: bent trigonal planar

b) hybridizations of C:  $sp^2$

b) Draw one Lewis structure of  $\text{XeF}_4$  in the space to the left and complete the table.

Lewis structure:



a) molecular geometry: square planar

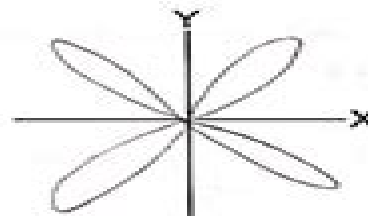
b) hybridizations of Xe:  $sp^3d^2$

5. Write the condensed electron configuration for the following: (2 marks)

a)  $\text{Tl}^+$   $[\text{Xe}] 6s^2 4f^{14} 5d^{10}$

b)  $\text{Ti}^{2+}$   $[\text{Ar}] 3d^2$

6. Draw the  $4d_{xy}$  orbital on the provided axes:



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