

Experiment 7  
 Irresistible Buffers?

Total Points: 22

Student's Name: Giselle Lawson

Lab Section: 604

The Effect of Acid on Solutions (6 pts)						
Beaker	Beakers 1-6					
	1	2	3	4	5	6
Contents	Dil H <sub>2</sub> O	50 mL NaCl	0.35 g NaAc + 50 mL HAc	0.60 g NaAc + 50 mL HAc	1.00 g NaAc + 50 mL HAc	2.00 g NaAc + 50 mL HAc
mL of HCl added	pH	pH	pH	pH	pH	pH
Black initial solvent, 0	6.68	6.28	4.56	4.75	5.02	5.36
1	1.87	1.78	4.31	4.57	4.85	5.22
2			4.02	4.37	4.72	5.11
3			3.63	4.17	4.58	5.02
4			2.58	3.94	4.45	4.94
5			1.76	3.65	4.32	4.84
6				3.03	4.16	4.75
7				1.91	4.01	4.66
8					3.83	4.59
9					3.60	4.59
10					3.24	4.42

1. (2 pts) Which beaker had the least amount of pH change? Explain why this occurred.

Beaker 6 had the least amount of pH change. This occurred because sodium acetate and acetic acid have the largest amount added to the beaker and together act a buffer system to the HCl. When HCl is added to the solution, the conjugate base in the buffer system consumes the hydronium ion, changing it to H<sub>2</sub>O and a weak acid. So as more HCl was added more of the conjugate base was being consumed leading to a lesser change in pH.

Beaker 6 had the least amount of pH change. This occurred because sodium acetate and acetic

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