

Sarah Kourba

Lab Partner: Peter Abanto

11/30/2017

Chem 201, Experiment 6

Electrochemical Cells

Procedure

See the pre-lab report on page 25 of my laboratory notebook for an outline of the general procedure. There were no changes made in the procedure.

Experimental Data & Calculated Results

	Concentration of Electrolyte (Red Lead)	Concentration of Electrolyte (Black Lead)	Experimental E_{cell} (V)	Calculated E_{cell} (V)	Percent Error
A	0.10M $ZnSO_4$	0.10M $CuSO_4$	-1.092	-1.100	0.72%
B	0.10M $ZnSO_4$	0.0010M $CuSO_4$	-1.010	-1.159	12.8%
C	0.0010 M $ZnSO_4$	0.10 $CuSO_4$	-1.034	-1.041	0.65%
D	0.10M $CuSO_4$	0.010 M $CuSO_4$	0.042	0.030	41.9%
E	0.10M $CuSO_4$	0.0010M $CuSO_4$	0.058	0.059	14.8%

Sample Calculations

- Refer to attached paper for sample calculations.

Discussion

For this experiment Cell C had the lowest percent error of 0.65% and Cell D had the highest percent error of 41.9%. My E_{cell} experimental values were close to that of the calculated E_{cell} calculated values. Cells A, B, C, and E are all below a percent error of 15% meaning that the values I had gotten from the experiment were not that far off from the calculated values. Cell D however, has a large percent error of 41.9%, meaning there is a significant difference between my experimental and calculated E_{cell} values. The purpose of percent error is to calculate how precise my experimental values

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