

In a precipitation reaction, two solutions containing soluble ionic salts are mixed. However, some of the ions, when the solutions are combined, can interact to form insoluble salts. When this occurs, a solid forms (i.e., the precipitant) and produces a cloudy solution and often falls to the bottom of the container. For example,

Solution A: Aqueous copper sulfate:  $\text{CuSO}_4 (\text{aq})$  or  $\text{Cu}^{2+} (\text{aq}) + \text{SO}_4^{2-} (\text{aq})$

Solution B: Aqueous sodium carbonate:  $\text{Na}_2\text{CO}_3 (\text{aq})$  or  $2 \text{Na}^+ (\text{aq}) + \text{CO}_3^{2-} (\text{aq})$

Notice that we can represent aqueous solutions (aq) in two ways, namely one with the ionic compound formula (aq) and one with the individual ions (aq). Both ways are equally valid. Also notice that when sodium carbonate goes into solution, we obtain one mole of carbonate anion for every two moles of sodium cations. Bulk solutions can have no net charge. Thus, when you sum the charges for all ions in solution, you must obtain zero. The integer number to the left of the sodium cation insures that the solution remains at zero net (or total) charge. Then, when the two solutions are mixed, we have the following chemical reaction equation:



The equation written above is called an ionic reaction equation. Notice that the sodium cation and the sulfate anion appear on both sides of the equation. This is because the combination  $\text{Na}_2\text{SO}_4$  is soluble. These ions are called spectator ions because they do not participate in the chemical reaction that yields the solid copper(II) carbonate. Thus, the net ionic reaction is



Notice how both of these reaction equations have the same number of each atom on the right hand and left hand side of the equation and that both of these equations also have the same net charge on both sides of the equation. Insuring that the charge and mass on both sides of a chemical reaction equation is the same are both part of balancing a chemical reaction equation.

## II. EXPERIMENT

- Given the following aqueous solutions, predict which mixtures of any two of these solutions would yield precipitants and determine the formula for the precipitant:
 

a. Copper(II) sulfate	b. Barium nitrate
c. Sodium chloride	d. Silver nitrate
e. Lead nitrate	f. Sodium sulfate
- Show your predictions to the instructor.
- Test to see if your predictions were correct by placing a small amount of the appropriate solution into a clean test tube and adding the second solution. You must test all of the reactions that you predict would form a precipitant. When the precipitant forms, record the color of the precipitant. You must also test at least two reactions that, based on Table 5.1, you predict would not form a precipitant.

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