

EXPERIMENT XI

ESTERIFICATION - The Preparation of Methyl Benzoate

REFERENCE: Introduction to Organic Laboratory Techniques, 3rd Ed., Pavia, Lampman and Kriz
Vacuum Distillation, pp. 558-562, 590-592

OBJECTIVES: The student should be able to:

1. Write up a step by step procedure for the reaction using information on the handout sheet, laboratory manuals and The Handbook of Chemistry and Physics.
2. Extract the reaction product by use of the properties of polarity and acidity.
3. Recover the unreacted benzoic acid.
4. Distill the product by vacuum distillation.
5. Calculate the amounts of reactants given the grams of product and the expected percent yield.
6. Identify the final product from physical properties, IR and NMR spectra.

main reaction equation

**REAGENTS:**

Reactants: Benzoic acid; methanol

Catalyst: conc. H₂SO₄, 4 ml

Calculate the amounts of the reactants required to produce 21.8 grams of methyl benzoate assuming a 80% yield.

Use this theoretical amount of benzoic acid.

Use a 518% excess of methanol.

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