

Balance elements that appear in only 1 compound on each side (C and H)



(4 Oxygens) (8 Oxygens)

Balance elements that appear in more than 1 compound on each side (O)



(2 Nitrogens) (6 Nitrogens)

Balance elements that appear in elemental form on either side last (N)



(3 Sulfur) (1 Sulfur)



(2 Gold) (1 Gold)



(2 Hydrogen) (6 Hydrogen)



If they do not add up correctly... use a fraction



(3 Carbon) (1 Carbon)



((6 Hydrogen) (2 Hydrogen)



(2 Oxygen) (7 Oxygens)



Multiply by 2 to make everything a whole number



Polyatomics are considered separately



(1 Sodium) (2 Sodium)



Mols and masses of reactants and products using stoichiometric factors (mol ratios)



This means that two moles of H₂ reacts with one mol of O₂ to generate 2 moles of H₂O

We can use this to create ratios that can be used in calculations

How many moles of H₂O can be obtained by 10.00 moles O₂ (with excess H₂) How many grams of H₂ do we need? 20 moles of H₂

2 mol H₂O/ 1 mol of O₂ is a conversion factor specific to the reaction, called the mol ratio

(10.00 moles O₂)(2 mol H₂O/1 mol O₂) = 20.00 mol H₂O produced

(20.00 mol H₂O)(18.02 g/mol) = 360.4 g H₂O

How many H₂ is required to react with all 10.00 mol of O₂?

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