

Pre lab n. 7

[a paragraph on the procedure of the current experiment]

First of all, we will need to prepare approximately 0.100M NaOH solution. We do this by adding 3 mL of stock solution of carbonate-free NaOH (approximately 19M) to 500 mL of distilled water (that has been heated up and cooled down). We then will standard the NaOH solution and then use the standardized NaOH solution to determine the pK_a of an unknown acid. At the end, we can calculate the concentration of the unknown acid using the volume of base at the equivalence point, its molarity, and the fact that we used 25 mL of acid.

Pre lab questions:

[write the answers, not the questions, of the prelab questions, see lab manual]

1. + A Bronsted-Lowry acid is any species that is capable of donating a proton H^+
+ A Bronsted-Lowry base is any species that is capable of accepting a proton, which requires a lone pair of electrons to bond to the H^+

2. $HA + H_2O \rightleftharpoons A^- + H_3O^+$

$$K_a = \frac{[H_3O^+][A^-]}{[HA]}$$

$$K_{eq} = \frac{[H_3O^+][A^-]}{[HA][H_2O]}$$

3. The pH at the equivalence point is not always equal to 7 in a neutralization titration because the conjugate base of a weak acid is weakly basic, the equivalence point of the titration reaches a pH above 7. Conversely, for the titration of a weak base with strong acid, the pH at the equivalence point is less than 7 because only the conjugate acid is present.
It will be 7 when it is the neutralization of a strong acid and a strong base.

4. $K_a = 6.5 \cdot 10^{-6} \Rightarrow pK_a = -\log(6.5 \cdot 10^{-6}) = 5.19$

5. Some current flow must occur which requires both a donor and acceptor for the electron.
6. A buffer is a solution that can resist pH change upon the addition of an acidic or basic components. It is able to neutralize small amounts of added acid or base, thus maintaining the pH of the solution relatively stable.

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