COLLEGE CHEMISTRY 1

Equilibrium

- Chemical equilibrium is the state in which the rates of the forward
- reverse reactions of a chemical reaction are equal.
- the equilibrium and how to calculate equilibrium constants for different reactions.
- Le Chatelier's principle to predict the effect of changes in concentration
- pressure, and temperature on the position of an equilibrium.

Acids and bases

- properties of acids and bases and their behavior in solution.
- calculate pH and pOH
- use acid-base titration to determine the concentration of an unknown acid or base.
- the concepts of buffers and how they maintain the pH of a solution.

Kimetics

- Chemical kinetics is the study of the rates of chemical reactions
- the factors that affect them.
- the principles of kinetics, including reaction rates, rate laws
- reaction mechanisms.
- experimental data to determine the rate law and reaction order for a chemical reaction.

Electrochemistry

- Electrochemistry is the study of the relationship between electricity and chemical reactions.
- the principles of electrochemistry, including oxidation-reduction reactions, electrochemical cells, and electrolysis.
- the Nernst equation to calculate the voltage of an electrochemical cell
- to use electrolysis to produce and analyze different chemical species.

Organic chemistry

- study of carbon-based compounds.
- the principles of bonding, functional groups, and reaction mechanisms in organic compounds.
- identify and name organic compounds
- to predict their behavior in different reactions.

Safety

- hazards associated with different chemicals and laboratory equipment
- to handle and dispose of chemicals safely.
- the principles of chemical hygiene and emergency procedures in case of accidents or spills.

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M Lipman

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Decoding Pkg Acp Cer College Chem 1 Lm Ch 101 052 S05: Revealing the Captivating Potential of Verbal Expression

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