

CHEM 1112 General Chemistry II Laboratory		
	Professors	Sections
	Kiran Tripathi, Ph.D	102, 103, 104, 113
Term	Spring 2024	
Meetings	Workshops: SLC 2.302, 3.102, 2.202, 3.102 Labs: SLC 1.205, 1.211, 3.210, 1.205	

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The easiest way to contact an instructor and/or TA is via e-mail.

Every instructor and TA will check their e-mail frequently and they try to respond as fast as possible.

Please always include both – your TA and your instructor – in your e-mail.

Office hours:

TA will post their booking hours on eLearning.

Mine office hours are on Thursday 1:30-3:30 pm. Room# **BE 3.330B**

General Course Information

Pre-requisites, Co-requisites, & other restrictions	Passing Grades in both CHEM 1311 and CHEM 1111 (General Chemistry I Lecture and Lab) or equivalents. No Audits allowed.
Course Description	This course is a continuation of CHEM 1111, it incorporates experiments in kinetics, acid base chemistry, chemical equilibrium, electrochemistry, and colligative properties.
Expected Learning Outcomes	<i>Students should be able to:</i> 1. Use graphing techniques and data analysis to evaluate data 2. Think critically through the analysis of experimental data 3. Determine the rate law of a chemical reaction 4. Determine equilibrium constant of a chemical system 5. Explain the effect of various parameters on equilibrium of a chemical system. 6. Generate and interpret pH titration curves
Required Texts & Materials	• Access to online labs through LabFlow. https://labflow.com. You will need to purchase an access code either online through their website or through the University Bookstore. Login information is posted on eLearning course page. Please make sure that you sign onto the correct lab section. • A non-programming calculator for calculations. • Access to eLearning. • A notebook for recording your data. • A combination code lock (shared with your lab group)

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