

CHEM 1211 Exam 5

Multiple Choice

Identify the choice that best completes the statement or answers the question.

Questions 1-30 are worth 6.67 points each for a total of 200 points.

1. Calcium hydroxide will neutralize phosphoric acid to produce solid calcium phosphate and water. Write and balance the chemical equation for this reaction. What is the sum of the coefficients?

- a. 4
- b. 11
- ☒ c. 12
- d. 16
- e. 28



Nitrogen gas reacts with hydrogen gas to produce gaseous ammonia:



Use the chemical equation to answer questions 2 through 5. The molecular weights of N_2 , H_2 , and NH_3 are 28.014 g/mol, 2.016 g/mol, and 17.031 g/mol.

2. How many moles of ammonia could be formed from the reaction of 3.25 moles of hydrogen gas with excess nitrogen gas?

- a. 1.63 moles
- ☒ b. 2.17 moles
- c. 3.25 moles
- d. 4.88 moles
- e. 6.50 moles

$$3.25 \text{ mol H}_2 \cdot \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} =$$

B

3. How many grams of nitrogen gas would be needed to completely react 4.52 grams of hydrogen gas?

- a. 0.976 grams
- b. 4.52 grams
- ☒ c. 20.9 grams
- d. 62.8 grams
- e. 188 grams

$$4.52 \text{ g H}_2 \cdot \frac{1 \text{ mol H}_2}{2.016 \text{ g H}_2} \cdot \frac{1 \text{ mol N}_2}{3 \text{ mol H}_2} \cdot \frac{28.014 \text{ g}}{1 \text{ mol N}_2}$$

C

4. How many grams of ammonia could be produced from the reaction of 5.23 grams of nitrogen with 5.23 grams of hydrogen?

- a. 0.445 grams
- b. 3.79 grams
- ☒ c. 7.58 grams
- d. 35.1 grams

$$5.23 \text{ g H}_2 \cdot \frac{1 \text{ mol H}_2}{2.016 \text{ g H}_2} \cdot \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \cdot \frac{17.031 \text{ g}}{1 \text{ mol}}$$

C

e) 52.6 grams

$$5.23 \text{ g N}_2 \cdot \frac{1 \text{ mol N}_2}{28.014 \text{ g N}_2} \cdot \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \cdot \frac{17.031 \text{ g}}{1 \text{ mol}}$$

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Table of Contents Pkgacp Chem 1211

1. Understanding the eBook Pkgacp Chem 1211
 - The Rise of Digital Reading Pkgacp Chem 1211
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkgacp Chem 1211
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkgacp Chem 1211
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkgacp Chem 1211
 - Personalized Recommendations
 - Pkgacp Chem 1211 User Reviews and Ratings
 - Pkgacp Chem 1211 and Bestseller Lists
5. Accessing Pkgacp Chem 1211 Free and Paid eBooks
 - Pkgacp Chem 1211 Public Domain eBooks
 - Pkgacp Chem 1211 eBook Subscription Services
 - Pkgacp Chem 1211 Budget-Friendly Options
6. Navigating Pkgacp Chem 1211 eBook Formats

- ePub, PDF, MOBI, and More
- Pkgacp Chem 1211 Compatibility with Devices
- Pkgacp Chem 1211 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pkgacp Chem 1211
 - Highlighting and Note-Taking Pkgacp Chem 1211
 - Interactive Elements Pkgacp Chem 1211
- 8. Staying Engaged with Pkgacp Chem 1211
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pkgacp Chem 1211
- 9. Balancing eBooks and Physical Books Pkgacp Chem 1211
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkgacp Chem 1211
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pkgacp Chem 1211
 - Setting Reading Goals Pkgacp Chem 1211
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pkgacp Chem 1211
 - Fact-Checking eBook Content of Pkgacp Chem 1211
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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