

## CH 131 – General Chemistry I

### Exam III – Review

1 atm = 760. torr = 760. mm Hg = 101.325 kPa = 1.01325 bar = 29.92 inches Hg = 14.7 psi

TABLE 10.2 Numerical Values of the Gas Constant  $R$  in Various Units

Units	Numerical Value
L-atm/mol-K	0.08206
J/mol-K*	8.314
cal/mol-K	1.987
m <sup>3</sup> -Pa/mol-K*	8.314
L-torr/mol-K	62.36

\*SI unit

- 1) Convert 60 mmHg to psi.
- 2) Write the corresponding relationships for Boyle's law, Charles's law, Avogadro's law, combined gas law and ideal gas law.
- 3) When you increase the temperature of a gas what will happen to the volume of the gas if  $P$  and  $n$  are constant?
- 4) If you increase the pressure of a gas what will happen to the volume of the gas under isothermal conditions?
- 5) A sample of gas initially has a volume of 859 mL at 565 K and 2.20 atm. What pressure will the sample have if the volume changes to 268 mL while the temperature is increased to 815 K?
- 6) What pressure will 14.0 g of CO exert in a 3.5 L container at 75°C?
- 7) What are the standard temperature and pressure for a gas?
- 8) What is the molar volume of a gas at STP?

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