

Chemistry 1011

Electromagnetic Radiation

Appreciate the relationship between frequency, wavelength and energy

- Wavelength (λ): the distance between two adjacent identical points of the wave
- Frequency (ν): the number of wave crests passing a given point per unit time
- All light waves travel at exactly the same speed – constant (c)
- When frequency increases, wavelength decreases, energy increases
- Atomic emission spectrum: atom supplied with energy, electromagnetic radiation emitted with fixed frequencies that are characteristic of the element

$$c = \lambda \nu$$

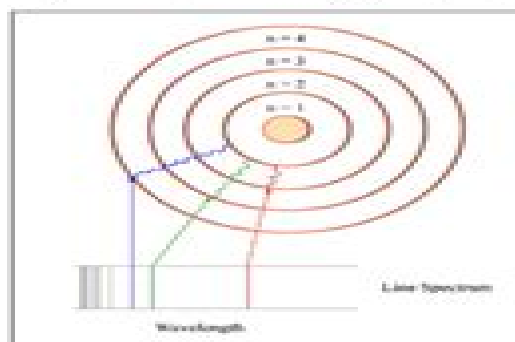
$$c = 2.998 \times 10^8 \text{ ms}^{-1}$$

Know that the Bohr model of an atom limits electrons to orbits of certain energies

- Electrons in atoms can only occupy certain energy levels (orbits)
- When an electron moves from one energy level to another, energy is absorbed or emitted – corresponds to light

Recognise the energy of light emitted from an atom represents the energy difference between two orbits

- The energy emitted corresponds to light of a specific energy/frequency



Know the maximum number of electrons the first, second and third shell

- First: 2

- Second: 8

- Third: 18

electrons the first, second and third shell

Electron configuration

Recognise the connection between an atom's electron configuration and its position in the periodic table

- The number of electrons is the atomic number
- The number of electrons in the outermost shell is the group number

Understand the limitations of the Bohr model of the atom

- Doesn't work well for atoms with multi-electron atoms

Recognise that electron shells contain sub-shells, and their connection to the Periodic Table organization

- Each shell is divided into sub-shells called s, p, d and f

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