

CHEM 145 Honors General Chemistry I

2021-08-17

Gas

| Kinetic Theory of Gas

Quantity	Unit	Definition
Pressure	N/m^2	$P = \frac{F}{A}$
Mole fraction of a	-	$X_a = \frac{n_a}{n_{total}}$
Partial pressure of a	N/m^2	$P_a = X_a P_{total}$
Pressure-volume work	J	$W = -P\Delta V$
Compressibility of gas	-	$z = \frac{PV}{nRT}$

Description	Equations
Ideal gas law	$PV = nRT$
Ideal gas constant and Boltzmann constant	$R = N_A k_B$
Average kinetic energy of one gas molecule	$\epsilon = \frac{3}{2} k_B T$
Root mean square speed gas	$v_{rms} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3k_B T}{m}}$
Mean speed of gas	$\bar{v} = \sqrt{\frac{8RT}{\pi M}} = \sqrt{\frac{8k_B T}{\pi m}}$
Most probable speed of gas	$v_{mp} = \sqrt{\frac{2RT}{M}} = \sqrt{\frac{2k_B T}{m}}$
Maxwell-Boltzmann speed distribution	$f(v) = 4\pi \left(\frac{m}{2\pi k_B T} \right)^{3/2} v^2 \exp \left(\frac{-mv^2}{2k_B T} \right)$
Van der Waals equation of state	$(P + a \frac{n^2}{V^2})(V - bn) = nRT$
Lennard-Jones potential	$V_{LJ} = 4\epsilon \left(\left(\frac{\sigma}{R} \right)^{12} - \left(\frac{\sigma}{R} \right)^6 \right)$

| Molecular Collisions and Rate Processes

Description	Equations
Molecule collision rate with wall	$Z_w \propto \frac{1}{4} \frac{N}{V} \bar{v} A = \frac{1}{4} \frac{N}{V} \sqrt{\frac{8RT}{\pi M}} A$
Graham's law of effusion	$\frac{\text{rate of effusion of A}}{\text{rate of effusion of B}} = \frac{N_A}{N_B} \sqrt{\frac{M_B}{M_A}}$
Molecule-molecule collision rate	$Z_1 = 4 \frac{N}{V} d^2 \sqrt{\frac{\pi RT}{M}}$
Mean free path	$\lambda = \frac{\bar{v}}{Z_1} = \frac{V}{\sqrt{2} \pi d^2 N}$

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