## Chem 132

Day 1 - Ch. 13

- Unsaturated solution not as much solute dissolved as possible (below the limit)
- Saturated Solution maximum amount of solute dissolved (at the limit)
- Supersaturated solution more than the maximum amount of solute (over the limit)
- Dilute solution little solute and lots of solvent
- Concentrated solution lots of solute in the given solvent
- Concentration = amount of solute / amount of solvent or solution
  - Solution = solute + solvent
    - Ag = solute + H2O
- Mass Percent = (mass of solute / mass of solution) x 100
- Bg NaOH are dissolved in 50g water. Calculate the mass percent.
  - (8g NaOH / (8g+50g)) x 100 = 13.8%
- What weight of KCI and water are needed to make 250g of a 5% KCI solution?
  - 5 = (g solute / 250 g solution) x 100 = 12.5g KCI.
  - 250 12.5 = 237.5 water
- Calculate the mole fraction of NH3 for a solution of 30.0g of NH3 in 70.0g H2O.
  - Mole fraction = mol A / total mol
  - 30 g NH3 / 17g = 1.76 mol NH3
  - 70 g water/ 18g = 3.89 mol water
    - 1.76 / (1.76+3.89) 0.312
- Molarity M moles of solute / Liters of solution
  - Molality m moles of solute / kg of solvent

H2O is always the solvent!

- You have a solution that is 20% by mass MgSO4 and the density of the solution is 1.30g/mL.
  - Density = mass / volume
  - 20g MgSO4 / 120g MgSO4 = 0.167 mol MgSO4
  - 80g = 0.08 kg water
  - Molality = 0.167 mol / 0.08kg = 2.09 m
  - 80g water / 18g = 4.44 mol H2O
  - Mole Fraction = 0.167 / (0.167 + 4.44) = 0.0360 MgSO4
  - 100g / 1.30g / 1000 = 0.0769L
  - Molarity = 0.167 / 0.0769 L = 2.17 M
- The density of a 1.099 M HCHO2 solution is 1.0115g/mL.
  - 1.099 mol HCHO2 x 46g = 50.6g HCHO2
  - 1 L x 1000 mL x 1.0115g = 1011.5g solution
  - Mass % = (50.6g / 1011.5g) × 100 = 5%
  - Molality = 1.099 / 0.9609 kg = 1.14 m

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