



# Best High Yield Savings Apple Watch Tricks

**S Ashworth**



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 $\text{H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g}) \rightarrow \Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow$



co<sub>2</sub> g 3 h 2 o g  $\Delta h$  283 kj g 1 c<sub>2</sub>h<sub>6</sub> g

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web thus for reaction 9 2 in which 1 mole of  $H_2SO_4$  and 1 mole of  $NaOH$  combine to form 2 moles of  $H_2O$   $\Delta H = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}}$

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