

pH Laboratory Activity

Name: _____

Station #	Test	Circle the Results	Circle the pH
A	Red Litmus	Red or Blue	Acid Base Neutral
	Blue Litmus	Red or Blue	
B	Red Litmus	Red or Blue	Acid Base Neutral
	Blue Litmus	Red or Blue	
C	Red Litmus	Red or Blue	Acid Base Neutral
	Blue Litmus	Red or Blue	

Station #	Test	Results	pH Value
A	Universal pH Paper	colour	pH
B	Universal pH Paper	colour	pH
C	Universal pH Paper	colour	pH

Station #	Test	pH Value
A	Normal Probe	pH
B	Normal Probe	pH

Station #4

Acid Base Neutralization Reaction of hydrogen chloride and sodium hydroxide via double displacement reaction.



Procedure

1. Add 10 drops of NaOH (base) on the evaporating dish.
2. Add 5 drops of phenolphthalein indicator on the evaporating dish and gently swirl the evaporating dish.
3. Slowly add 5 drops of HCl (acid) on the evaporating. Swirl the evaporating dish.
4. Slowly add another 5 drops of HCl (acid) on the evaporating. Swirl the evaporating dish. When the colour disappears, the solution has been neutralized.

Note: If the solutions are the exact same concentration, 10 drops of acid should neutralize 10 drops of the base to produce water and sodium chloride.

5. How many drops did you need to neutralize the sodium hydroxide (base)?

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RS Peters



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Table of Contents Ph Laboratory And Clabroom Activities1999spiralbound

1. Understanding the eBook Ph Laboratory And Clabroom Activities1999spiralbound
 - The Rise of Digital Reading Ph Laboratory And Clabroom Activities1999spiralbound
 - Advantages of eBooks Over Traditional Books
2. Identifying Ph Laboratory And Clabroom Activities1999spiralbound
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ph Laboratory And Clabroom Activities1999spiralbound
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ph Laboratory And Clabroom Activities1999spiralbound
 - Personalized Recommendations
 - Ph Laboratory And Clabroom Activities1999spiralbound User Reviews and Ratings

- Ph Laboratory And Clabroom Activities1999spiralbound and Bestseller Lists
- 5. Accessing Ph Laboratory And Clabroom Activities1999spiralbound Free and Paid eBooks
 - Ph Laboratory And Clabroom Activities1999spiralbound Public Domain eBooks
 - Ph Laboratory And Clabroom Activities1999spiralbound eBook Subscription Services
 - Ph Laboratory And Clabroom Activities1999spiralbound Budget-Friendly Options
- 6. Navigating Ph Laboratory And Clabroom Activities1999spiralbound eBook Formats
 - ePub, PDF, MOBI, and More
 - Ph Laboratory And Clabroom Activities1999spiralbound Compatibility with Devices
 - Ph Laboratory And Clabroom Activities1999spiralbound Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ph Laboratory And Clabroom Activities1999spiralbound
 - Highlighting and Note-Taking Ph Laboratory And Clabroom Activities1999spiralbound
 - Interactive Elements Ph Laboratory And Clabroom Activities1999spiralbound
- 8. Staying Engaged with Ph Laboratory And Clabroom Activities1999spiralbound
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ph Laboratory And Clabroom Activities1999spiralbound
- 9. Balancing eBooks and Physical Books Ph Laboratory And Clabroom Activities1999spiralbound
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ph Laboratory And Clabroom Activities1999spiralbound
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ph Laboratory And Clabroom Activities1999spiralbound
 - Setting Reading Goals Ph Laboratory And Clabroom Activities1999spiralbound
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ph Laboratory And Clabroom Activities1999spiralbound
 - Fact-Checking eBook Content of Ph Laboratory And Clabroom Activities1999spiralbound
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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


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