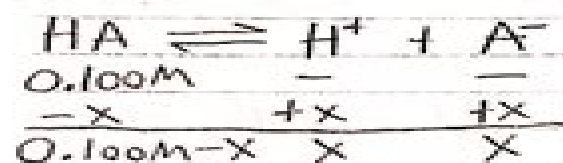


- Calculate K_a for the following acids:
- 1) 100.0 mL of 0.100 M HA with a pH of 3.40

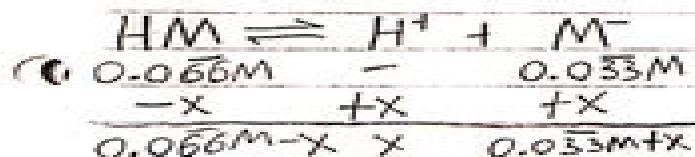


$$\begin{aligned}
 \text{pH} &= 3.40 \rightarrow \text{pH} = -\log[\text{H}^+] \\
 3.40 &= -\log(x) \rightarrow x = 3.98 \times 10^{-4}
 \end{aligned}$$

$$K_a = \left(\frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \right) = \left(\frac{(x)^2}{(0.100\text{M}-x)} \right)$$

$$K_a = \left(\frac{(3.98 \times 10^{-4})^2}{(0.100\text{M} - 3.98 \times 10^{-4})} \right) \rightarrow \boxed{K_a = 1.59 \times 10^{-6}}$$

- 2) 100.0 mL of 0.100 M HM is mixed with 50.0 mL of 0.100 M NaM and has a pH of 5.50



$$\begin{aligned}
 \text{pH} &= 5.50 \rightarrow \text{pH} = -\log[\text{H}^+] \\
 5.50 &= -\log(x) \rightarrow x = 3.16 \times 10^{-6}
 \end{aligned}$$

Dilution Problem Due to HM and NaM

$$\begin{aligned}
 (C_1)(V_1) &= (C_2)(V_2) \\
 (0.100\text{M HM})(100.0\text{mL}) &= (C_2)(150.0\text{mL}) \\
 (C_2) &= 0.066\text{M HM}
 \end{aligned}$$

Dilution Problem Due to NaM and HM

$$\begin{aligned}
 (C_1)(V_1) &= (C_2)(V_2) \\
 (0.100\text{M NaM})(50.0\text{mL}) &= (C_2)(150.0\text{mL}) \\
 (C_2) &= 0.033\text{M NaM}
 \end{aligned}$$

$$K_a = \left(\frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \right) = \left(\frac{(x)(0.033\text{M}+x)}{(0.066\text{M}-x)} \right)$$

$$K_a = \left(\frac{(3.16 \times 10^{-6})(0.033\text{M} + 3.16 \times 10^{-6})}{(0.066\text{M} - 3.16 \times 10^{-6})} \right) \rightarrow K_a = 1.58 \times 10^{-6}$$

Problems And Solutions For College Chemistry And General Chemistry

**William Franklin Kieffer, Robert K.
Fitzgerel**



Problems And Solutions For College Chemistry And General Chemistry:

Problems and Solutions Supplement to College Chemistry and General Chemistry, Seventh Editions Henry Fuller Holtzclaw, William Harrison Nebergall, William R. Robinson, 1984 Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields **Schaum's Outline of College Chemistry, Ninth Edition** Jerome Rosenberg, Lawrence Epstein, Peter Krieger, 2009-10-16 Confusing Textbooks Missed Lectures Not Enough Time Fortunately for you there s Schaum s Outlines More than 40 million students have trusted Schaum s to help them succeed in the classroom and on exams Schaum s is the key to faster learning and higher grades in every subject Each Outline presents all the essential course information in an easy to follow topic by topic format You also get hundreds of examples solved problems and practice exercises to test your skills This Schaum s Outline gives you Practice problems with full explanations that reinforce knowledge Coverage of the most up to date developments in your course field In depth review of practices and applications Fully compatible with your classroom text Schaum s highlights all the important facts you need to know Use Schaum s to shorten your study time and get your best test scores Schaum s Outlines Problem Solved *General Chemistry* Henry Fuller Holtzclaw, William Harrison Nebergall, William R. Robinson, 1984 **1968, Teil 1b: Cephaline-Harnalyse** , 2021-12-20 Keine ausführliche Beschreibung f r 1968 Teil 1b Cephaline Harnalyse verfügbar 1967, Teil 1b: Ch17-Haranalyse Helga Völz, Christian Weiske, 2022-02-07 Keine ausführliche Beschreibung f r 1967 Teil 1b Ch17 Haranalyse verfügbar *Biology/science Materials* Carolina Biological Supply Company, 1991 **Publishers' Circular and General Record of British and Foreign Literature, and Booksellers' Record** , 1896 College Chemistry William Harrison Nebergall, Henry Fuller Holtzclaw, William R. Robinson, 1980 THEORY and PROBLEMS of COLLEGE CHEMISTRY , 1966 **The Chemical News and Journal of**

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