

When light shines on a metal, an electron can absorb the light and use the energy to break free from the metal. It becomes a free electron, called a photoelectron.

Wave Model Prediction:

Time delay: dim light would take some time to eject electrons (needs time for the energy to build up). The kinetic energy of the photoelectrons should depend on the intensity of the light. This is because in the wave model, the energy of the light is related to the intensity. So bright light has lots of energy and should give the photoelectrons lots of energy. Changing the frequency should have no effect on the observations. Intense light should be able to eject electrons, no matter what the frequency is. **Observations:**

The kinetic energy of the photoelectrons was the same for both dim and bright light of a particular frequency. The kinetic energy of the photoelectrons depended on the frequency of the light. E.g., electrons ejected by violet light moved faster than electrons emitted by green light. If the frequency was too low, then no photoelectrons were ejected. (The minimum frequency needed to eject electrons is called the **threshold frequency** f_0). Above the threshold frequency, electrons were ejected **immediately**, even for extremely dim light. There was no time delay between turning on the light and having electrons emitted.

Particle Model Explain:

Work Function ϕ : The minimum amount of energy needed to remove an electron from a particular metal. **Stopping voltage V_s :** The minimum amount of voltage needed to repel the most energetic photoelectrons from reaching the collector electrode. **Photon:** a packet (particle) of light. **Photoelectron:** an electron that has been freed from a metal by absorbing light.

FORCE: $F = eV$
 λ of e^- in an e^- gun =

(SYNCHROTRON)
 $|$ (BOLD = $h \cdot m \cdot v$)
(if $E = \text{Joules}$) (if $E = eV$)

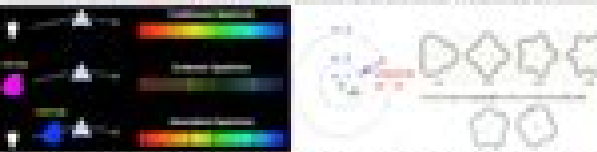
Wave: Two-slit interference pattern. (Interference is a wave phenomenon – particles cannot interfere); Diffraction; Polarisation **Particle:** The photoelectric effect.



Gradient = Planck's constant
x-intercept = threshold frequency
y-intercept = work function
Louis de Broglie suggested that matter should also exhibit both particle and wave, with the same formula for wavelength as for light:

if you send individual photons through a double slit, they can be detected on a screen as single particles. But if you let the experiment build up over time, an interference pattern forms. **Energy:**

Momentum:
Light is made up of particles called photons. The energy of each photon is determined by the frequency of the light: $E = hf$. Low frequencies will not eject electrons because energy is related to frequency ($E = hf$). **Intensity** doesn't matter because photons per second can be adjusted. The **wavelength** will not affect the energy of each individual electron since each electron interacts with **individual** only one photon. Even very dim light can eject electrons because you only need one photon to hit one electron. There is no time delay because spectrum can cover a **small** number of **photon** waves. You don't have to wait for energy to build up.



Spectra explained by Bohr model of the atom: Electrons orbit the nucleus in particular energy levels (shells). Electrons can only be in these energy levels; they cannot be between them. The energy levels are different for different elements. Electrons can move between levels by absorbing or emitting photons (particle model of light). The energy of the photon must exactly match the difference between the energy levels: $\Delta E = hf$. Louis de Broglie explained that matter (electrons) could have wave-like properties. He said that the electron could only maintain a steady energy level (orbit) if it was a standing wave. So, the only orbits allowed are when the circumference is exactly equal to a whole number multiple of the electron wavelength.

Heisenberg's Uncertainty Principle: The more accurately you know the position of an object, the less accurately you know its momentum, and vice-versa. It is a fundamental limit to the possible accuracy of any physical measurement. E.g. uncertainty in the y position affects the uncertainty in the y momentum.

Electromagnetic Radiation (visible light):

Electrons will emit light when they move from a higher energy level to a lower energy level. This process produces **discrete** light (quantised), such as that seen in an emission spectrum. Charged particles will produce light **whenever they accelerate**. As acceleration is a continuous quantity, the light produced is **continuous**. Any frequency of light can be produced.

Sources of light:

Incandescent lights: Continuous spectrum, **incoherent**.
Light emitted by hot objects: Continuous spectrum, **incoherent**.
Lasers: **Monochromatic**, **coherent**, (in phase)
Synchrotrons: Continuous spectrum, **incoherent**.
LEDs: **Monochromatic**, **incoherent**.

Significant diffraction occurs when the wavelength is greater to or equal to the size of the gap or obstacle.

So to the wave before double slit the wave particles density of electrons can be used to replace the quantised energy levels of the atom. (7 marks)

- Electrons have wave-like properties and hence form a wavefront.
- Electrons will only pass through a double-slit when exactly lined up with the screen. (This is where the circumference of the slit is a whole number multiple of the wavelength).
- Therefore, only specific wavelengths and their associated energies would be allowed.



Explain why the electron microscope gives a clearer image than the optical microscope. (4 marks)

- The clarity of the image is related to the definition of the most diffraction orders the image forms.
- The waves of diffraction is proportional to $\frac{1}{\lambda}$.
- Electrons have a much smaller wavelength than for optical microscope light. Therefore they diffract less and therefore create a clearer image.

The wavelengths of visible light ranges between 400 and 700nm, hence the ratio is much greater than 1 before it is the length scale of the virus particles, significant diffraction of light occurs contributing to the inability to distinguish the edges of objects of this scale. The wavelengths of X-rays are in the range of 1×10^{-11} m whereas that molecules are in the range of 1×10^{-9} m. Therefore, the ratio is much less than 1, with little diffraction occurring, it is possible to distinguish the edge of objects on the scale of DNA.



Project Physics Unit 4 Light Electroma

Ayodele O. Ogunleye



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1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Project Physics Unit 4 Light Electroma
- Chapter 2: Essential Elements of Project Physics Unit 4 Light Electroma
- Chapter 3: Project Physics Unit 4 Light Electroma in Everyday Life
- Chapter 4: Project Physics Unit 4 Light Electroma in Specific Contexts
- Chapter 5: Conclusion

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6. In chapter 5, this book will draw a conclusion about Project Physics Unit 4 Light Electroma. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Project Physics Unit 4 Light Electroma.

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