

CHM-113 PRACTICE QUIZ 6 KEY
CHAPTER 11

1. Identify the state of matter for A, B, and C.

A) particles are close together in random positions with about equal kinetic energy and intermolecular forces.

B) particles are close together in fixed positions with low kinetic energy

C) particles are far apart with greater kinetic energy and low intermolecular forces

	A	B	C
a)	solid	liquid	gas
b)	gas	liquid	solid
c)	liquid	solid	gas
d)	liquid	gas	solid

2. Rank the states of matter in order of *increasing* kinetic energy.

a) solid < liquid < gas

b) gas < liquid < solid

c) solid = liquid < gas

d) solid < liquid = gas

3. Rank the states of matter in order of *decreasing* intermolecular forces.

a) gas > liquid = solid

b) solid = liquid > gas

c) solid > liquid > gas

d) gas > liquid > solid

4. When a substance in its liquid state undergoes a decrease in kinetic energy _____.

a) it may melt

b) the intermolecular forces will decrease

c) it may solidify

d) it may evaporate

5. When a substance changes from a liquid to a solid, which of the following is TRUE?

a) The molecules of the substance get heavier.

b) The molecules of the substance change shape.

c) The molecules of the substance move further away from each other.

d) The molecules of the substance connect more strongly to one another.

6. Which causes the high density of solids?

a) The particles are more massive than those in liquids.

b) The intermolecular forces between particles are weak.

c) The particles are packed closely together.

d) The kinetic energy of the particles is very high.

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