

# Experiment 10: Galvanic Cells

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Report: Galvanic Cells

Table 1

Cell Description	Ag	Cu	Zn	Mg	Pb
Ag	X	0.50V	1.10V	1.70V	0.80V
Cu	0.50V	X	0.70V	1.40V	0.40V
Zn	1.10V	0.70V	X	0.90V	0.30V
Mg	1.70V	1.40V	0.90V	X	0.60V
Pb	0.80V	0.40V	0.30V	0.60V	X

Table 2 Reduction half-reactions and standard potentials

Cell	Reduction half-reaction	Potential
Ag	$\text{Ag}^+ + e^- \rightarrow \text{Ag}$	0.80V
Cu	$\text{Cu}^{2+} + 2e^- \rightarrow \text{Cu}$	0.34V
Zn	$\text{Zn}^{2+} + 2e^- \rightarrow \text{Zn}$	-0.76V
Mg	$\text{Mg}^{2+} + 2e^- \rightarrow \text{Mg}$	-1.10V
Pb	$\text{Pb}^{2+} + 2e^- \rightarrow \text{Pb}$	-0.13V

## Discussion:

The reduction potential of a half-cell is a measure of the tendency of a species to be reduced. The more positive the reduction potential, the more likely the species is to be reduced. In this experiment, we measured the reduction potentials of several half-cells and compared them to the standard reduction potentials listed in Table 2.

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