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Christian Drosten



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web q1 arrange the bonds in each of the following sets in order of increasing polarity c f be f o f o cl s br c p c s b f n o q2 determine based on electronegativity and symmetry if the following molecules are polar or nonpolar hcn h c n cf4 c f 4 cocl2 c o c l 2 nf3 n f 3 q3

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web when the atoms in a bond are not the same the bonding electrons are shared and the bond is electronegativities equally

unequally non polar polar unequally polar the degree of polarity of a bond between any two atoms is determined by consulting a table of

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web the relative electronegativities of the bonded atoms is h c cl and so the bond moments all point toward the cl end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as BF_3 trigonal planar CH_4 tetrahedral PF_5 trigonal bipyramidal and SF_6

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