

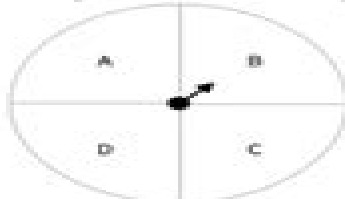
## Making Predictions With Probability

Page 1 of 2

scribble space

1

The spinner below has 4 equal sections.



If you spin it 256 times, how many times would you expect to spin *A* OR *B*?

*Answer*

2

You roll a 6-sided die 252 times.



How many times would you expect to roll a 3?

*Answer*

3

A bakery recently sold a number of cupcakes. 6 cupcakes were vanilla, 10 were strawberry, and 8 were chocolate. Using experimental probability, predict how many of the next 120 cupcakes sold at the bakery will be vanilla cupcakes.

# Math Worksheet Grade Oscar Predictions This Week

**G Orfield**



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The Times-picayune Index ,1992      **Journal of Educational Psychology** ,1928      **The Continuing Study of**  
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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$$

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web start with 8 3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8 3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149 4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137 1 \text{ g } \text{H}_2\text{O}}{149 4 \text{ g } \text{H}_2\text{O}} \times 100 = 91 77\%$  problems 6 8 are similar to problem 5

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calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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g  $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$  c determine the percent yield  
yield actual  $\times 100 = 260.0 \text{ g} \times 100 / 86.01 = 302.3 \text{ g}$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be  
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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s  
remains when the reaction is complete

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 $\text{CH}_3\text{OH}$  can be produced by the following reaction  $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of co is  
reacted with 8.6 g of  $\text{H}_2$

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