



# NAMELESS WORLD

A Fantasy World Map

# Google Maps Fantasy Football Ideas

**H Kauffman**



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web 1 consider the reaction  $\text{i}_2\text{O}_5 \text{ g } 5 \text{ co g } 5 \text{ co}_2 \text{ g } \text{i}_2 \text{ g a } 80.0 \text{ grams of iodine v oxide } \text{i}_2\text{O}_5 \text{ reacts with } 28.0 \text{ grams of carbon monoxide co determine the mass of iodine i}_2 \text{ which could be produced b if in the above situation only } 0.160 \text{ moles of iodine i}_2 \text{ was produced i what mass of iodine was produced}$

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web limiting reactant theoretical and percentage yields key  $4 \text{ ko}_2 \text{ s } 2 \text{ h}_2 \text{ o l } 4 \text{ koh aq } 3 \text{ o}_2 \text{ g if a reaction vessel contains } 0 \text{ mol ko}_2 \text{ and } 0 \text{ mol h}_2 \text{ o what is the limiting reactant}$

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  

$$8 3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149 4 \text{ g } \text{H}_2\text{O}$$
 is your theoretical yield so percent yield  

$$\frac{137 1 \text{ g } \text{H}_2\text{O}}{149 4 \text{ g } \text{H}_2\text{O}} \times 100 = 91 77\%$$
 problems 6 8 are similar to problem 5

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web 1 csun edu hcchm001 introchemhandouts html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5 0 g of n what is the limiting reagent if 8 52 g are actually formed what is the

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
 if the reaction of 21 3 grams of  $\text{O}_2$  produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have  

$$\frac{21 3 \text{ g } \text{O}_2}{32 \text{ g } \text{O}_2} = 0 66 \text{ moles}$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160 g  $\text{Al}$  x 1 mol  $\text{Al}$  x 2 mol  $\text{Al}_2\text{O}_3$  x 101 964 g  $\text{Al}_2\text{O}_3$  302 3 g  $\text{Al}_2\text{O}_3$  26 982 g  $\text{Al}$  4 mol  $\text{Al}$  1 mol  $\text{Al}_2\text{O}_3$  c determine the percent yield  
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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web may 20 2018 learning objectives identify the limiting reactant limiting reagent in a given chemical reaction calculate how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is complete

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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$   $\times$  1333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage  
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